

REVIEW

4

SECTION 4.2

Ionic and Covalent Bonding

1. **Explain** why solid table salt is a poor conductor of electricity, while salt water is a good conductor of electricity.

2. **Explain** why table salt does not melt easily.

3. **Contrast** ionic and covalent bonds.

4. **Explain** why a triple bond between two nitrogen atoms is stronger than a double bond between two oxygen atoms.

5. **Explain** how it is possible for a compound to have both ionic and covalent bonds.

6. **Predict** whether a gold ring would be a good conductor of electricity. What kind of bonds does gold have? How do these bonds explain gold's properties?
