

## REVIEW

## 4

## SECTION 4.3

# Compound Names and Formulas

1. **Explain** the difference between iron(II) nitrate and iron(III) nitrate. What is the significance of the Roman numerals?

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2. **Name** the following ionic compounds, keeping in mind that a transition metal cation must include its charge.

_____	a. $\text{TiO}_2$
_____	b. $\text{BaCl}_2$
_____	c. $\text{CuCl}_3$
_____	d. $\text{KI}$
_____	e. $\text{SrCl}_2$
_____	f. $\text{CuBr}_2$

3. **Describe** how covalent compounds are named.

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4. **Write** the chemical formulas for the following covalent compounds:

_____	a. dinitrogen pentoxide
_____	b. carbon monoxide
_____	c. carbon tetrachloride
_____	d. nitrogen trifluoride
_____	e. phosphorus pentachloride
_____	f. disulfur dichloride

5. **Contrast** molecular formulas and empirical formulas.

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